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pH= 4.76 + log (4.04/5.96) = 4.59. Thus, a change in pH of 0.19 pH units will be observed. The buffering capacity = (change in OH-) / (change in pH)= 1 mM/0.19 pH units=5.26 mmol/pH unit. If we had added 10 mL of 0.1 M HCl instead of NaOH, after the addition, acetic acid = 7.96 mmol and acetate = 2.04 mmol.

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